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Field Book of Practical Mineralogy; How to Examine and Report on Mines, Designed for the Use of Prospectors, Mining Men, Engineers and Others

By George Washington Miller

Rarebooksclub.com, United States, 2013. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can usually download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1901 edition. Excerpt: .yield the odor of phosphuretted hydrogen. IN THE WET WAY. Orthophosphates.--1. Magnesia mixture precipitates white $MgNH_4PO_4$. 2. $AgNO_3$ precipitates light-yellow Ag_3PO_4 , soluble in HNO_3 and NH_4OH . 3. $(NH_4)_2MoO_4 + HNO_3$ precipitates yellow ammonium phospho-molybdate; composition variable. The precipitate is soluble in NH_4OH , in excess of phosphoric acid, and is prevented by organic substances, such as tartaric acid. Pyrophosphate.---1. $MgSO_4$ precipitates white $Mg_3P_2O_8$, soluble in an excess of either solution. $NI(OH)_3$ fails to precipitate it from these solutions. On boiling it separates again. By this reaction pyro can be detected in the presence of phosphoric acid. 2. $(NH_4)_2MoO_4 + HNO_3$ does not give a precipitate until orthophosphate is formed. Most of the pyrophosphates of the heavy metals (Ag an exception) are soluble in alkali pyrophosphates (distinction from orthophosphates). 3. $AgNO_3$ precipitates white $Ag_3P_2O_8$, soluble in HNO_3 and NH_4OH . Addition of an alkali aids the precipitation. Metaphosphoric Acid.--1....



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