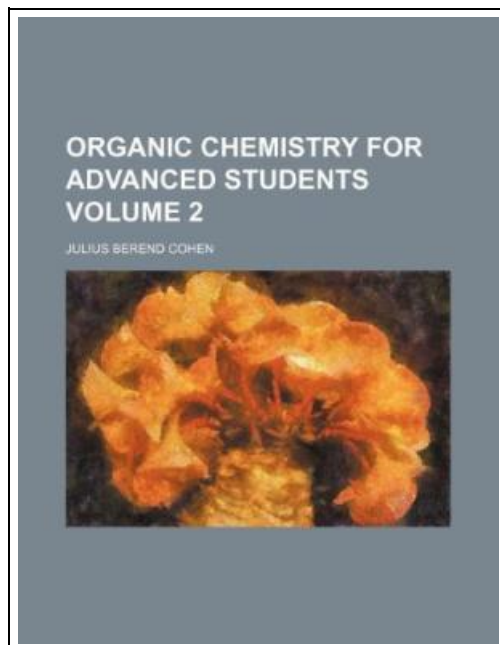


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Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1913 Excerpt: .stages and is therefore a composite reaction; cf. Orton and King, Trans. Chem. Soc., 1911, 89, 1369; also Orton and Jones, Trans. Chem. Soc., 1909, 95, 1456. $\text{CH}_3\text{CO}_2\text{C}_2\text{H}_5 + \text{NaOH} = \text{CH}_3\text{CO}_2\text{Na} + \text{C}_2\text{H}_5\text{OH}$ The saponification proceeds at a rate which can be conveniently measured at temperatures between 0 and 25 if dilute solutions are employed. If the original solution contains a gram.-mols. (mole) of ester and a mols of hydroxide per unit volume, and if x mols of ester have been saponified after time t , the concentrations of the reacting substances at this moment will be $a-x$ and $b-x$ respectively. According to the mass law, the speed of the change will be given by $\frac{dx}{dt} = k_2(a-x)(b-x)$ and this on integration becomes, $\int \frac{dx}{(a-x)(b-x)} = k_2 \int \frac{1}{(a-x)(b-x)} dt$ In the following table are given the data obtained by Reicher for the saponification of ethyl acetate at 15-8, the alkali being present in excess ($b > a$) in the one experiment, whilst the ester predominated in the second ($a > b$). The quantities of saponified ester (x) are expressed in terms of the standard acid solution which was used in following the progress of the change. If the reacting substances are present in equivalent proportions ($a = b$), the rate of change at time t is given by $\frac{dx}{dt} = k(a-x)^2$ from which $\int \frac{1}{(a-x)^2} dx = k \int \frac{1}{(a-x)^2} dt$ That this is in agreement with the actual course of saponification under these conditions is shown by the...



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