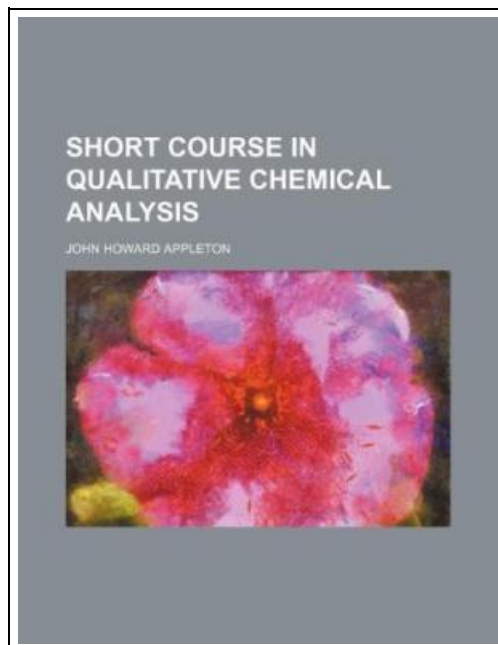


Short Course in Qualitative Chemical Analysis



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Reviews

This publication is definitely not simple to begin on studying but quite fun to see. It really is full of knowledge and wisdom I am just effortlessly can get a satisfaction of studying a created pdf.

(Alfreda Bradtke)

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Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1895 Excerpt: .Potassic nitrate (KN03) certainly contains no Nitric acid (HN0s); but it contains N03, which is a group of elements contained in Nitric acid, and which group is the representative of Nitric acid. But it has arisen--from the inconvenience attending the use of such terms as representative, acid radicle, etc.--that chemists commonly speak of salts as containing acids. Thus, sulphates are spoken of as containing Sulphuric acid, phosphates as containing Phosphoric acid, and so on. 129. The more common acid radicles are divided into six Groups. In testing for acid radicles, no strictly continuous processes of isolation from a single portion of substance can be used; most of the tests are performed upon separate portions of the original substance, set aside fm this use. FIRST GROUP OF ACIDS. THE DEFLAGRATING ACIDS. 130. This group includes those acids whose salts readily deflagrate, when heated on charcoal. It consists of--Nitric acid, H N Os. Chloric acid, H Cl O, . Preliminary Tests. 131. Heat upon charcoal a portion of the dry salt. (If the substance is not a salt, evaporate a portion of the solution to dryness, and test the residue.) If deflagration ensues, it indicates the probable presence of a Nitrate or a Chlorate. The deflagration is rapid combustion of the coal, due to the fact that the salts, in question, contain large amounts of oxygen, held by feeble affinity. iKNO, + C (heated) = K, CO, + N2 O, . 2KCIO, + 3 C (heated) = 2...



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